

CRAYON

Cathode

- Reduction

Selectively
discharge lower
reactive cations

Anode

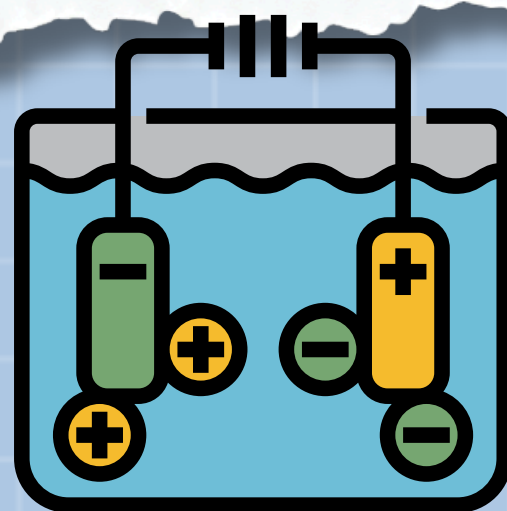
- Oxidation

Selectively
discharge ...

1. If anode is reactive (ie not made of graphite, carbon or platinum), it oxidizes itself

2. If anode is inert and electrolyte is diluted/aqueous, higher [Anion] is selectively discharged

3. If anode is inert and electrolyte is aqueous, OH⁻ has a greater ease of discharge than halides and complex ions



SIMPLE CELL

-0.10V

1. Electrons flow from more reactive to less reactive metal

ie. electrons flow from Na to Pt

2. **Magnitude** of voltage is determined by differences in reactivity

ie. Na/Pt has a larger magnitude of voltage compared to Na/Cu simple cell

3. **Polarity** of voltage is determined by direction of flow of electrons

ie Na/Pt for -0.10 V thus Pt/Na will be for +0.10 V

